

REACTIONS OF GLYCALs WITH XENON FLUORIDE: AN IMPROVED SYNTHESIS OF  
2-DEOXY-2-FLUOROSACCHARIDES

W. Korytnyk\* and S. Valentekovic-Horvat

Department of Experimental Therapeutics, Grace Cancer Drug Center,

Roswell Park Memorial Institute, Buffalo, New York 14263

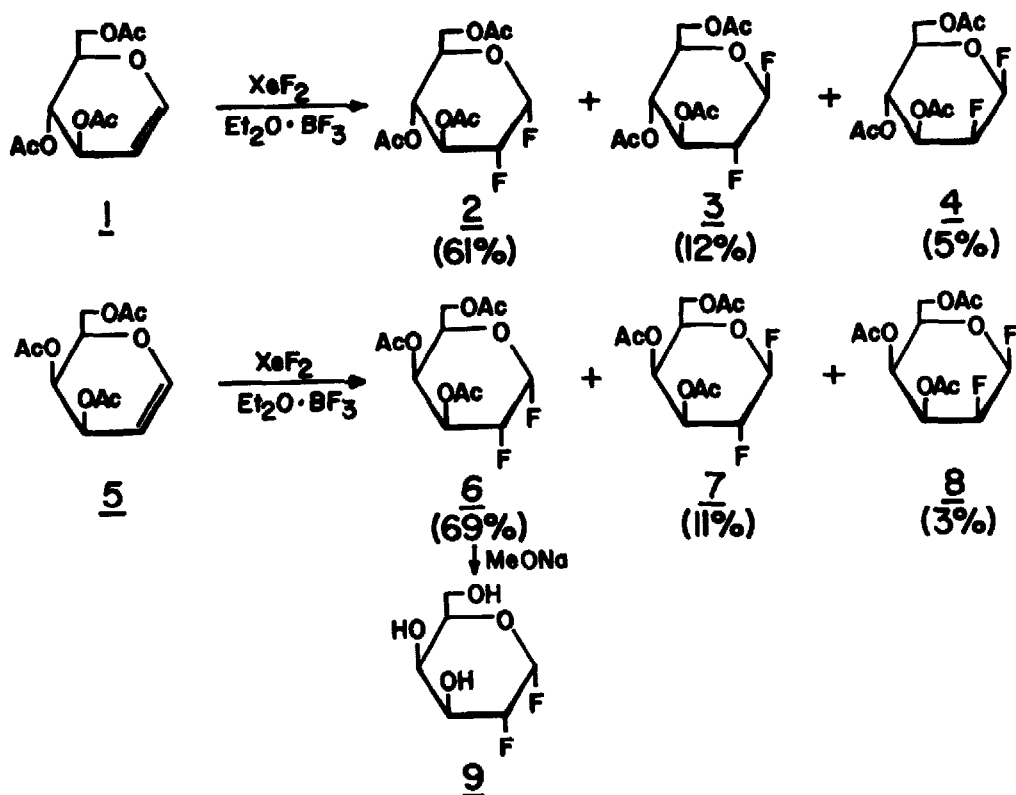
Summary: The 1,2-double bond in acetylated glycal<sub>s</sub> has been fluorinated with XeF<sub>2</sub> in the presence of BF<sub>3</sub> to give 1,2-deoxy-1,2-difluorosaccharides. A mechanism for this reaction has been proposed. This method represents an improvement in the synthesis of 2-deoxy-2-fluorosaccharides.

The synthesis and biological activities of fluorinated carbohydrates has received considerable attention;<sup>1,2</sup> of particular interest are 2-deoxy-2-fluorohexoses, such as 2-deoxy-2-fluoro-D-glucopyranose, both in regards to their biological activities as inhibitors or modifiers of cell surface glycoconjugates<sup>3</sup> as well as for emission tomography, if labeled with <sup>18</sup>F.<sup>4</sup>

The synthesis of 2-deoxy-2-fluoro-D-glucose has been achieved by conventional displacement reactions<sup>5</sup>, by the electrophilic additions of trifluoromethyl hypofluoride<sup>6</sup> or elemental fluorine<sup>7</sup> to 1 and subsequent hydrolysis. Although these electrophilic addition reactions simplify the synthesis, they involve the handling of toxic and corrosive gases. In addition, CF<sub>3</sub>OF has often led to complex mixtures of products<sup>8</sup> and its commercial availability has been restricted.

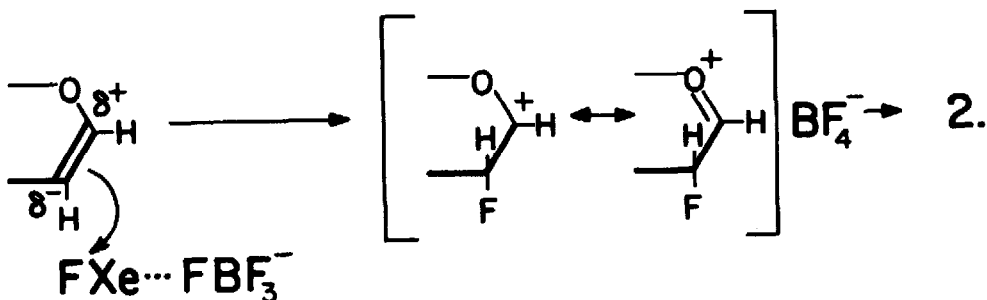
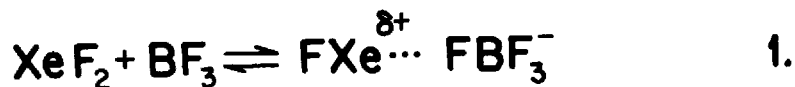
We have adapted xenon difluoride<sup>9</sup>, an easily handled and available solid for fluorinations of glycal<sub>s</sub>. This is exemplified by reactions with 3,4,6-tri-O-acetyl-1,5-anhydro-2-deoxy-D-arabino-hex-1-enitol (3,4,6-tri-O-acetyl-D-glucal) 1 and the D-lyxo epimer (tri-O-acetyl-D-galactal) 5 in Scheme 1 (yields are given in parenthesis):

### Scheme 1: Fluorinations of Glycals with Xenon Fluoride



In a typical experiment, **1** (1 mmol) was dissolved in  $\text{Et}_2\text{O}$  (5ml) and added while stirring to  $\text{XeF}_2$  (1 mmole) at room temperature. A solution of  $\text{BF}_3 \cdot \text{OEt}_2$  (0.2 mmole) in dry benzene (5ml) was added dropwise and the stirring was continued for 24 hrs, when no starting material could be detected. Products **2** to **4** have been separated on a silica gel column and their identity confirmed by physical data reported in the literature<sup>10</sup> and by analysis of their  $^1\text{H}$  and  $^{19}\text{F}$  NMR spectra. The fluorination of **5** under identical conditions gave products that were stereochemically similar to that obtained from **1**, indicating that an opposite configuration of the acetoxy group in the 4 position has little effect on the stereochemical course of the reaction. Hydrolysis with acid of either **2** or **3** gave 2-deoxy-2-fluoro-D-glucose and of **6** or **7** 2-deoxy-2-fluoro-D-galactose, respectively. The deacetylation of the adducts has also been carried with  $\text{NaOMe}$  yielding almost quantitatively 2-deoxy-2-fluoro-D-glycopyranosyl fluorides, as shown by the synthesis of **9** from **6** (Scheme 1).

The mechanism of the fluorination reaction can be rationalized by the following reaction sequence:



In the first reaction  $\text{XeF}_2$  forms a complex with  $\text{BF}_3$ ; this is followed by the nucleophilic attack of the  $\pi$ -electron system of the 1,2-double bond upon the fluorine of the complex to afford the 2-fluorinated and resonance stabilized carbocation, initially complexed with fluoroborate ion (reaction 2). In the last step (reaction 3),  $\text{BF}_4^-$  donates  $\text{F}^-$  to the anomeric carbon. The initial attack of the complex occurs predominantly from the less hindered side of the double bond resulting in a certain degree of stereoselectivity.

#### Acknowledgements

This study was supported in part by research grant CA-08793 and Grace Cancer Drug Center Core Grant CA-24538 from the National Institutes of Health. The NMR facility has received partial support from the Institute Core Grant CA-16056. We thank Ms. Onda Dodson Simmons for determining the NMR spectra.

References

1. A.B. Foster and J.H. Westwood, Pure Appl. Chem., 35, 147 (1973).
2. (a) P.W. Kent, in "Ciba Foundation Symposium on Carbon-Fluorine Compounds", Elsevier, Associated Publishers, New York, 1972, p. 169.  
(b) J.E.G. Barnett, *idem*, p. 95.
3. R. Datema and R.T. Schwarz, Biochem. J., 184, 113 (1979).
4. T. Ido, C.-N. Wan, V. Casella, J.S. Fowler, A.P. Wolf, M. Reivich, and D.E. Kuhl, J. Labelled Compounds and Radiopharmaceuticals, 14, 175 (1978).
5. J. Pacak, J. Podesva, Z. Tocik and M. Cerny, Collection Czechoslovak Chem. Communication, 37, 2589 (1972).
6. J. Adamson, A.B. Foster, L.D. Hall, R.N. Johnson, and R.H. Hesse, Carbohydrate Res., 15, 351 (1970).
7. T. Ido, C.-N. Wan, J.S. Fowler, and A.P. Wolf, J. Org. Chem., 42, 2341 (1977).
8. (a) P.W. Kent and S.D. Dimitrijevic, J. Fluorine Chem., 10, 455 (1977); (b) G.C. Butchard and P.W. Kent, Tetrahedron, 35, 2439 (1979).
9. A. Gregorcic and M. Zupan, J. Org. Chem., 44, 1255, 4120 (1979).
10. J. Adamson, A.B. Foster, L.D. Hall, and R.H. Hesse, Chem. Commun. 309 (1969).

(Received in USA 4 January 1980)